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In physical chemistry, you use physics and math to understand and describe atoms and molecules, getting to the theoretical basis for thermodynamics, kinetics, electron configurations, and even the phases of matter.

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$\ln K = \left( \frac{\Delta S^\circ}{R} \right) - \left( \frac{\Delta H^\circ}{RT} \right)$  The equilibrium for the dissociation of one mole of iodine vapour is  $3.94 \times 10^{-4}$  at 900K and the standard enthalpy is 154 kJmol<sup>-1</sup>. Calculate the equilibrium constant at 1200K assuming the standard enthalpy remains constant.  $\ln (K_2/K_1) = - \frac{\Delta H^\circ}{R} \left( \frac{1}{T_2} - \frac{1}{T_1} \right)$ , ANS =  $6.7 \times 10^{-2}$  at 1200K.

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### Physical Chemistry - American Chemical Society

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Often these courses provide a broad survey of the concepts and in-depth treatment of selected topics. The challenge of designing a one-semester course in physical chemistry is to determine the important principles that govern the physical and chemical behavior of matter within the context of the course emphasis.

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