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8-3 In the formation of
a negative ion, a
neutral atom gains
one or more

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electrons. Again, note that in the neutral atom the number of protons equals the number of electrons. However, the ion contains more electrons than protons, making this over-all charge on this ion negative. e

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Ionic Compounds

Answer Key 8.2 Ionic

Formulas Compounds

are Neutral They have

no net charge so you

must have enough

cations and anions to

equal zero. Na^+ and

Cl^- make NaCl since

$+1$ and $-1 = 0$ $+ \text{Na}$

and CN^- make NaCN

since $+1$ and $-1 = 0$ $2+$

Ba and Cl^- make

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BaCl₂

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8. In a crystal lattice of an ionic compound,

- a. ions of a given charge are clustered together, far from ions of the opposite charge.
- b. ions are surrounded by ions of the opposite charge.
- c. a sea of electrons surrounds the ions.
- d.

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that holds two atoms
together is called a(n)
answer Chemical
bond. question Such
an attachment may
form by the attraction
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CHAPTER 8

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Applying Scientific
Methods A college
chemistry student is
studying the

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properties of four unknown compounds, W, X, Y and Z. She has been informed that one of them is ionic and that the other three are covalent. Of the latter, she has been told that the attractions between the formula units are dispersion forces in one case, hydrogen bonds in another, and

Download Free Chapter 8 Ionic Covalent (network) bonds in another.

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CHAPTER 8

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Applying S ...

Section 8.2 What is an ionic bond? In your textbook, read about forming ionic bonds and the

characteristics of ionic

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Compounds. Circle
the letter of the choice
that best completes
the statement or
answers the question.

1. An ionic bond is a.
attraction of an atom
for its electrons. b.
attraction of atoms for
electrons they share.

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