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Play this game to review Chemical Reactions. Using the following equation: $\text{Fe}_2\text{O}_3(\text{s}) + 3\text{H}_2(\text{g}) \rightarrow 2\text{Fe}(\text{s}) + 3\text{H}_2\text{O}(\text{l})$ How many moles of Fe can be made from 6 moles H_2 ? (This is a one step conversion using mole ratio)

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Quiz ... This quiz will give you some more practice in solving the various kinds of stoichiometric calculations. Remember that you cannot solve the questions without a balanced chemical equation (none will ...

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Stoichiometry Stoichiometry is the practice of predicting the amount of product or reactant in a chemical

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equation based on a known amount of one of the other products or reactants. The ability to do these calculations is the culmination of all of the basic skills learned throughout the first semester.

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Stoichiometry: Mixed Problems (KEY)

1) $N_2 + 3H_2 \rightarrow 2NH_3$ What volume of NH_3 at STP is produced if 25.0 of N_2 is reacted with an excess of H_2 ? $3 \ 3 \ 3 \ 2 \ 3$
 $2 \ 2 \ 2 \ 40.0L \ NH \ 1 \ mol \ NH \ 22.4L \ NH$
 $1 \ mol \ N \ 2 \ mol \ NH \ 28.0g \ N \ 25.0g \ N$
 $1 \ mol \ N \times \times \times = 2) \ 2KClO_3 \rightarrow 2KCl +$

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302 If 5.0g of KClO_3 is decomposed, what volume of O_2 is produced at STP? 2

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Mixed Stoichiometry Problems . 1.

$2\text{H}_2 + \text{O}_2 \rightarrow 2\text{H}_2\text{O}$. a). How many moles of H_2 would be required to produce

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5.0 moles of water? 5.0 moles water.

b). What mass of H₂O is formed when H₂ reacts with 384 g of O₂? 432g H₂.

2. H₂SO₄ + 2NaOH (Na₂SO₄ + 2H₂O.

a). Balance this equation. Look above.

b).

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Learn how to use mole ratios derived from balanced chemical equations to calculate amounts of substances consumed and produced in chemical reactions.

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of conservation of mass where the total mass of the Page 11/21

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